



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 99

1. $\text{Cu} + \text{O}_2 = \text{Cu}_2\text{O}$
2. $\text{LiOH} + \text{H}_2\text{SO}_4 = \text{Li}_2\text{SO}_4 + \text{H}_2\text{O}$
3. $2 \text{NH}_3 + \text{CuO} = \text{Cu} + \text{N}_2 + 3 \text{H}_2\text{O}$
4. $\text{Ca}(\text{OH})_2 + \text{CH}_4 = \text{CaC}_2 + \text{H}_2 + 2 \text{H}_2\text{O}$
5. $2 \text{NH}_3 + \text{CuO} = \text{Cu} + \text{N}_2 + 3 \text{H}_2\text{O}$
6. $\text{H}_3\text{PO}_3 = \text{PH}_3 + \text{H}_3\text{PO}_4$
7. $\text{B}_2\text{H}_6 + \text{H}_2\text{O} = 2 \text{H}_3\text{BO}_3 + \text{H}_2$
8. $\text{FeS}_2 + 11 \text{O}_2 = \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
9. $6 \text{NaH} + \text{BF}_3 = \text{B}_2\text{H}_6 + \text{NaF}$
10. $\text{SF}_4 + \text{H}_2\text{O} = \text{H}_2\text{SO}_3 + \text{HF}$
11. $4 \text{OCl}^{-1} + \text{S}_2\text{O}_3^{-2} + \text{OH}^{-1} = \text{SO}_4^{-2} + 4 \text{Cl}^{-1} + \text{H}_2\text{O}$
12. $\text{HS} + 5 \text{O}_2 = \text{SO}_2 + 2 \text{H}_2\text{O}$
13. $2 \text{KBr} + \text{PbO}_2 + \text{HNO}_3 = \text{Pb}(\text{NO}_3)_2 + \text{Br}_2 + \text{KNO}_3 + 2 \text{H}_2\text{O}$
14. $\text{Al} + \text{Fe}_3\text{N}_2 = \text{AlN} + 3 \text{Fe}$
15. $\text{C}_1\text{H}_5\text{O}_2 + 5 \text{O}_2 = \text{CO}_2 + 10 \text{H}_2\text{O}$
16. $\text{PCl}_5 + \text{H}_2\text{O} = \text{H}_3\text{PO}_4 + \text{HCl}$
17. $\text{Ca}(\text{NO}_3)_2 + 2 \text{K}_3\text{PO}_4 = \text{Ca}_3(\text{PO}_4)_2 + \text{KNO}_3$
18. $\text{Bi}(\text{OH})_3(\text{s}) + \text{HNO}_3(\text{aq}) = \text{Bi}(\text{NO}_3)_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$
19. $2 \text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
20. $\text{H}_2\text{SO}_4 + \text{K}_2\text{O} = \text{KHSO}_4 + \text{H}_2\text{O}$



ANSWERS

1. $4 \text{ Cu} + \text{O}_2 = 2 \text{ Cu}_2\text{O}$
2. $2 \text{ LiOH} + \text{H}_2\text{SO}_4 = \text{Li}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$
3. $2 \text{ NH}_3 + 3 \text{ CuO} = 3 \text{ Cu} + \text{N}_2 + 3 \text{ H}_2\text{O}$
4. $\text{Ca}(\text{OH})_2 + 2 \text{ CH}_4 = \text{CaC}_2 + 3 \text{ H}_2 + 2 \text{ H}_2\text{O}$
5. $2 \text{ NH}_3 + 3 \text{ CuO} = 3 \text{ Cu} + \text{N}_2 + 3 \text{ H}_2\text{O}$
6. $4 \text{ H}_3\text{PO}_3 = \text{PH}_3 + 3 \text{ H}_3\text{PO}_4$
7. $\text{B}_2\text{H}_6 + 6 \text{ H}_2\text{O} = 2 \text{ H}_3\text{BO}_3 + 6 \text{ H}_2$
8. $4 \text{ FeS}_2 + 11 \text{ O}_2 = 2 \text{ Fe}_2\text{O}_3 + 8 \text{ SO}_2$
9. $6 \text{ NaH} + 2 \text{ BF}_3 = \text{B}_2\text{H}_6 + 6 \text{ NaF}$
10. $\text{SF}_4 + 3 \text{ H}_2\text{O} = \text{H}_2\text{SO}_3 + 4 \text{ HF}$
11. $4 \text{ OCl}^{\{-\}} + \text{S}_2\text{O}_3^{\{-2\}} + 2 \text{ OH}^{\{-\}} = 2 \text{ SO}_4^{\{-2\}} + 4 \text{ Cl}^{\{-\}} + \text{H}_2\text{O}$
12. $4 \text{ HS} + 5 \text{ O}_2 = 4 \text{ SO}_2 + 2 \text{ H}_2\text{O}$
13. $2 \text{ KBr} + \text{PbO}_2 + 4 \text{ HNO}_3 = \text{Pb}(\text{NO}_3)_2 + \text{Br}_2 + 2 \text{ KNO}_3 + 2 \text{ H}_2\text{O}$
14. $2 \text{ Al} + \text{Fe}_3\text{N}_2 = 2 \text{ AlN} + 3 \text{ Fe}$
15. $4 \text{ C}_1\text{H}_5\text{O}_2 + 5 \text{ O}_2 = 4 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
16. $\text{PCl}_5 + 4 \text{ H}_2\text{O} = \text{H}_3\text{PO}_4 + 5 \text{ HCl}$
17. $3 \text{ Ca}(\text{NO}_3)_2 + 2 \text{ K}_3\text{PO}_4 = \text{Ca}_3(\text{PO}_4)_2 + 6 \text{ KNO}_3$
18. $\text{Bi}(\text{OH})_3(\text{s}) + 3 \text{ HNO}_3(\text{aq}) = \text{Bi}(\text{NO}_3)_3(\text{aq}) + 3 \text{ H}_2\text{O}(\text{l})$
19. $2 \text{ Al}(\text{OH})_3 + 3 \text{ H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 6 \text{ H}_2\text{O}$
20. $2 \text{ H}_2\text{SO}_4 + \text{K}_2\text{O} = 2 \text{ KHSO}_4 + \text{H}_2\text{O}$



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