



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 89

- $\text{CdCl}_2 + \text{NaOH} = \text{Cd(OH)}_2 + \text{NaCl}$
- $\text{HS} + 5 \text{O}_2 = 4 \text{SO}_2 + \text{H}_2\text{O}$
- $3 \text{TiO}_2 + \text{Al} = 3 \text{Ti} + \text{Al}_2\text{O}_3$
- $\text{Li} + \text{CO}_2 + 6 \text{H}_2\text{O} = \text{LiHCO}_3 + 5 \text{H}_2$
- $\text{Bi}_2\text{O}_3 + \text{NaClO} + 2 \text{NaOH} = 2 \text{NaBiO}_3 + \text{NaCl} + \text{H}_2\text{O}$
- $\text{BaCl}_2(\text{aq}) + 2 \text{H}_3\text{PO}_4(\text{aq}) = \text{Ba}_3(\text{PO}_4)_2(\text{s}) + \text{HCl}(\text{aq})$
- $\text{H}_3\text{SbO}_4 + \text{Na}_2\text{O} = \text{NaH}_2\text{SbO}_4 + \text{H}_2\text{O}$
- $\text{KOH} + \text{H}_3\text{PO}_4 = \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- $2 \text{NiO(OH)} + 3 \text{Cd} + \text{H}_2\text{O} = 2 \text{Ni(OH)}_2 + \text{Cd(OH)}_2$
- $\text{HClO}_4 + \text{Ba(OH)}_2 = \text{H}_2\text{O} + \text{Ba(ClO}_4)_2$
- $5 \text{FeCl}_2 + \text{KMnO}_4 + \text{HCl} = 5 \text{FeCl}_3 + \text{MnCl}_2 + \text{H}_2\text{O} + \text{KCl}$
- $2 \text{SO}_2 + \text{HNO}_3 = \text{H}_2\text{SO}_4 + \text{NO}$
- $\text{Fe}_2\text{O}_3 + \text{HCl} = 2 \text{FeCl}_3 + \text{H}_2\text{O}$
- $\text{Pb(NO}_3)_2 = 2 \text{PbO} + \text{NO}_2 + \text{O}_2$
- $\text{Cu(OH)}_2 + \text{HNO}_3 = \text{Cu(NO}_3)_2 + \text{H}_2\text{O}$
- $\text{KO}_2 + 2 \text{H}_2\text{O} = \text{O}_2 + 4 \text{KOH}$
- $\text{Cr}_2\text{O}_7^{2-} + 427 \text{S}_2^{2-} + 432 \text{H}^{(+)} = \text{Cr}_3^{(+)} + 854 \text{S} + 216 \text{H}_2\text{O}$
- $\text{Fe} + 10 \text{O}_2 + \text{H}_2\text{O} = \text{Fe(OH)}_2$
- $\text{Zn} + \text{AgCl} = \text{ZnCl}_2 + \text{Ag}$
- $\text{C}_{20}\text{H}_{42} + 61 \text{O}_2 = \text{CO}_2 + 42 \text{H}_2\text{O}$



ANSWERS

1. $\text{CdCl}_2 + 2 \text{NaOH} = \text{Cd(OH)}_2 + 2 \text{NaCl}$
2. $4 \text{HS} + 5 \text{O}_2 = 4 \text{SO}_2 + 2 \text{H}_2\text{O}$
3. $3 \text{TiO}_2 + 4 \text{Al} = 3 \text{Ti} + 2 \text{Al}_2\text{O}_3$
4. $2 \text{Li} + \text{CO}_2 + 6 \text{H}_2\text{O} = 2 \text{LiHCO}_3 + 5 \text{H}_2$
5. $\text{Bi}_2\text{O}_3 + 2 \text{NaClO} + 2 \text{NaOH} = 2 \text{NaBiO}_3 + 2 \text{NaCl} + \text{H}_2\text{O}$
6. $3 \text{BaCl}_2(\text{aq}) + 2 \text{H}_3\text{PO}_4(\text{aq}) = \text{Ba}_3(\text{PO}_4)_2(\text{s}) + 6 \text{HCl}(\text{aq})$
7. $2 \text{H}_3\text{SbO}_4 + \text{Na}_2\text{O} = 2 \text{NaH}_2\text{SbO}_4 + \text{H}_2\text{O}$
8. $3 \text{KOH} + \text{H}_3\text{PO}_4 = \text{K}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
9. $2 \text{NiO(OH)} + 3 \text{Cd} + 4 \text{H}_2\text{O} = 2 \text{Ni(OH)}_2 + 3 \text{Cd(OH)}_2$
10. $2 \text{HClO}_4 + \text{Ba(OH)}_2 = 2 \text{H}_2\text{O} + \text{Ba(ClO}_4)_2$
11. $5 \text{FeCl}_2 + \text{KMnO}_4 + 8 \text{HCl} = 5 \text{FeCl}_3 + \text{MnCl}_2 + 4 \text{H}_2\text{O} + \text{KCl}$
12. $2 \text{SO}_2 + 2 \text{HNO}_3 = \text{H}_2\text{SO}_4 + 6 \text{NO}$
13. $\text{Fe}_2\text{O}_3 + 6 \text{HCl} = 2 \text{FeCl}_3 + 3 \text{H}_2\text{O}$
14. $2 \text{Pb(NO}_3)_2 = 2 \text{PbO} + 4 \text{NO}_2 + \text{O}_2$
15. $\text{Cu(OH)}_2 + 2 \text{HNO}_3 = \text{Cu(NO}_3)_2 + 2 \text{H}_2\text{O}$
16. $4 \text{KO}_2 + 2 \text{H}_2\text{O} = 3 \text{O}_2 + 4 \text{KOH}$
17. $3 \text{Cr}_2\text{O}_7^{\{-\}} + 427 \text{S}_2^{\{-\}} + 432 \text{H}^{\{+\}} = 2 \text{Cr}_3^{\{+\}} + 854 \text{S} + 216 \text{H}_2\text{O}$
18. $10 \text{Fe} + 10 \text{O}_2 + \text{H}_2\text{O} = 10 \text{Fe(OH)}_2$
19. $\text{Zn} + 2 \text{AgCl} = \text{ZnCl}_2 + 2 \text{Ag}$
20. $2 \text{C}_{20}\text{H}_{42} + 61 \text{O}_2 = 40 \text{CO}_2 + 42 \text{H}_2\text{O}$



Thanks for downloading our free printable.

We have thousands of such resources in our blog for teachers and parents.

[You can download them for free here!](#)

Free Printables from Go Science Girls – Fair Usage Policy

You can ...

- Download and save this free printable from gosciencegirls.com to your computer.
- Print this file and use it as many times as you want in your home, classrooms or for your library.
- Feel free to link our blog post where your visitors can find and download this printable for free.
- When you post online about this resource – please give due credit to gosciencegirls.com

You Cannot ...

- Access this file or download it from other sites apart from gosciencegirls.com
- Other websites cannot link to this pdf directly. If required, they are welcomed to link to the blog post from where this pdf can be downloaded.
- The ownership of this pdf rests with GoScienceGirls. No one can claim ownership for this file.
- You are not allowed to sell printed copies of this file to others.
- You are not allowed to store this file electronically and redistribute it (only personal use is allowed).

Further Questions?

Feel free to email us at contactgosciencegirls@gmail.com for any further questions and suggestions. We would love to hear from you. We promise to respond back as soon as we can.