



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 80

1. $\text{K} + 2 \text{H}_2\text{O} = \text{KOH} + \text{H}_2$
2. $2 \text{CaH}_2 + \text{H}_2\text{O} = 2 \text{CaOH} + \text{H}_2$
3. $4 \text{Fe} + \text{SnCl}_4 = 4 \text{FeCl}_3 + \text{Sn}$
4. $\text{K}_2\text{Cr}_2\text{O}_7 + \text{HCl} = \text{KCl} + 2 \text{CrCl}_3 + 7 \text{H}_2\text{O} + 6 \text{Cl}$
5. $\text{N}_2\text{O}_5 = \text{N}_2\text{O}_4 + \text{O}_2$
6. $\text{HCl} + \text{Fe}_2\text{O}_3 = 2 \text{FeCl}_3 + \text{H}_2\text{O}$
7. $2 \text{CrBr}_3 + \text{Na}_2\text{SiO}_3 = \text{Cr}_2(\text{SiO}_3)_3 + \text{NaBr}$
8. $\text{Al}(\text{NO}_3)_3 + 3 \text{H}_2\text{SO}_4 = \text{HNO}_3 + \text{Al}_2(\text{SO}_4)_3$
9. $\text{L}_2 + 10 \text{HNO}_3 = 6 \text{HLO}_3 + \text{NO} + 2 \text{H}_2\text{O}$
10. $\text{FeCl}_2 + \text{KMnO}_4 + 8 \text{HCl} = \text{FeCl}_3 + \text{MnCl}_2 + 4 \text{H}_2\text{O} + \text{KCl}$
11. $8 \text{KI} + \text{H}_2\text{SO}_4 = 4 \text{K}_2\text{SO}_4 + 4 \text{I}_2 + \text{H}_2\text{S} + \text{H}_2\text{O}$
12. $2 \text{C}_6\text{H}_6(\text{l}) + \text{O}_2(\text{g}) = 12 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
13. $\text{Ba}(\text{OH})_2 + 2 \text{H}_6\text{TeO}_6 = \text{Ba}_3(\text{H}_3\text{TeO}_6)_2 + \text{H}_2\text{O}$
14. $2 \text{Ca}_2(\text{PO}_4)_2 + \text{SiO}_2 + 12 \text{C} = \text{CaSiO}_3 + \text{P}_4 + 12 \text{CO}$
15. $\text{K}_2\text{Cr}_2\text{O}_7 + \text{HCl} = 2 \text{KCl} + 2 \text{CrCl}_3 + \text{Cl}_2 + 7 \text{H}_2\text{O}$
16. $3 \text{Cl}_2 + \text{H}_2\text{O} + 6 \text{AgNO}_3 = 5 \text{AgCl} + \text{HNO}_3 + \text{AgClO}_3$
17. $\text{CN}^{\{-\}} + \text{MnO}_4 = \text{CNO}^{\{-\}} + \text{MnO}_2$
18. $2 \text{C}_7\text{H}_{10}\text{N} + \text{O}_2 = 14 \text{CO}_2 + \text{H}_2\text{O} + 2 \text{NO}_2$
19. $\text{C}_2\text{O}_4 + \text{H}_2\text{O} = \text{H}_2\text{C}_2\text{O}_4 + \text{OH}$
20. $3 \text{Cl}_2 + \text{H}_2\text{O} = \text{HCl} + \text{HClO}_3$



ANSWERS

1. $2 \text{ K} + 2 \text{ H}_2\text{O} = 2 \text{ KOH} + \text{H}_2$
2. $2 \text{ CaH}_2 + 2 \text{ H}_2\text{O} = 2 \text{ CaOH} + 3 \text{ H}_2$
3. $4 \text{ Fe} + 3 \text{ SnCl}_4 = 4 \text{ FeCl}_3 + 3 \text{ Sn}$
4. $\text{K}_2\text{Cr}_2\text{O}_7 + 14 \text{ HCl} = 2 \text{ KCl} + 2 \text{ CrCl}_3 + 7 \text{ H}_2\text{O} + 6 \text{ Cl}$
5. $2 \text{ N}_2\text{O}_5 = 2 \text{ N}_2\text{O}_4 + \text{O}_2$
6. $6 \text{ HCl} + \text{Fe}_2\text{O}_3 = 2 \text{ FeCl}_3 + 3 \text{ H}_2\text{O}$
7. $2 \text{ CrBr}_3 + 3 \text{ Na}_2\text{SiO}_3 = \text{Cr}_2(\text{SiO}_3)_3 + 6 \text{ NaBr}$
8. $2 \text{ Al}(\text{NO}_3)_3 + 3 \text{ H}_2\text{SO}_4 = 6 \text{ HNO}_3 + \text{Al}_2(\text{SO}_4)_3$
9. $3 \text{ L}_2 + 10 \text{ HNO}_3 = 6 \text{ HLO}_3 + 10 \text{ NO} + 2 \text{ H}_2\text{O}$
10. $5 \text{ FeCl}_2 + \text{KMnO}_4 + 8 \text{ HCl} = 5 \text{ FeCl}_3 + \text{MnCl}_2 + 4 \text{ H}_2\text{O} + \text{KCl}$
11. $8 \text{ KI} + 5 \text{ H}_2\text{SO}_4 = 4 \text{ K}_2\text{SO}_4 + 4 \text{ I}_2 + \text{H}_2\text{S} + 4 \text{ H}_2\text{O}$
12. $2 \text{ C}_6\text{H}_6(\text{l}) + 15 \text{ O}_2(\text{g}) = 12 \text{ CO}_2(\text{g}) + 6 \text{ H}_2\text{O}(\text{g})$
13. $3 \text{ Ba}(\text{OH})_2 + 2 \text{ H}_6\text{TeO}_6 = \text{Ba}_3(\text{H}_3\text{TeO}_6)_2 + 6 \text{ H}_2\text{O}$
14. $2 \text{ Ca}_2(\text{PO}_4)_2 + 4 \text{ SiO}_2 + 12 \text{ C} = 4 \text{ CaSiO}_3 + \text{P}_4 + 12 \text{ CO}$
15. $\text{K}_2\text{Cr}_2\text{O}_7 + 14 \text{ HCl} = 2 \text{ KCl} + 2 \text{ CrCl}_3 + 3 \text{ Cl}_2 + 7 \text{ H}_2\text{O}$
16. $3 \text{ Cl}_2 + 3 \text{ H}_2\text{O} + 6 \text{ AgNO}_3 = 5 \text{ AgCl} + 6 \text{ HNO}_3 + \text{AgClO}_3$
17. $2 \text{ CN}^{\{-\}} + \text{MnO}_4 = 2 \text{ CNO}^{\{-\}} + \text{MnO}_2$
18. $2 \text{ C}_7\text{H}_{10}\text{N} + 21 \text{ O}_2 = 14 \text{ CO}_2 + 10 \text{ H}_2\text{O} + 2 \text{ NO}_2$
19. $\text{C}_2\text{O}_4 + 2 \text{ H}_2\text{O} = \text{H}_2\text{C}_2\text{O}_4 + 2 \text{ OH}$
20. $3 \text{ Cl}_2 + 3 \text{ H}_2\text{O} = 5 \text{ HCl} + \text{HClO}_3$



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