



## BALANCE THE GIVEN CHEMICAL EQUATIONS

### Worksheet - 74

- $\text{Pb(s)} + \text{ \_\_\_\_ AgNO}_3(\text{aq}) = \text{Pb(NO}_3)_2 + \text{ \_\_\_\_ Ag(s)}$
- $\text{Fe}_2\text{O}_3 + \text{ \_\_\_\_ CO} = 3 \text{ CO}_2 + \text{ \_\_\_\_ Fe}$
- $2 \text{ Al(s)} + 2 \text{ NaOH(aq)} + \text{ \_\_\_\_ H}_2\text{O(l)} = \text{ \_\_\_\_ Na(Al(OH)}_4\text{)(aq)} + 3 \text{ H}_2(\text{g})$
- $2 \text{ KMnO}_4 + \text{ \_\_\_\_ HCl} = 2 \text{ KCl} + 2 \text{ MnCl}_2 + 8 \text{ H}_2\text{O} + \text{ \_\_\_\_ Cl}_2$
- $\text{C}_6\text{H}_{12}\text{O}_6 + \text{ \_\_\_\_ O}_2 = 6 \text{ CO}_2 + \text{ \_\_\_\_ H}_2\text{O}$
- $\text{ \_\_\_\_ Zn} + 8 \text{ HNO}_3 = 3 \text{ Zn(NO}_3)_2 + 2 \text{ NO} + \text{ \_\_\_\_ H}_2\text{O}$
- $\text{ \_\_\_\_ Fe}^{\{2+\}} + \text{MNO}_4^{\{-\}} + 8 \text{ H}^{\{+\}} = \text{ \_\_\_\_ Fe}^{\{3+\}} + \text{MN}^{\{2+\}} + 4 \text{ H}_2\text{O}$
- $\text{ \_\_\_\_ HCl} + \text{K}_2\text{SO}_3 = \text{ \_\_\_\_ KCl} + \text{H}_2\text{O} + \text{SO}_2$
- $\text{TiCl}_4 + \text{ \_\_\_\_ H}_2\text{O} = \text{TiO}_2 + \text{ \_\_\_\_ HCl}$
- $2 \text{ HSbCl}_4 + \text{ \_\_\_\_ H}_2\text{S} = \text{Sb}_2\text{S}_3 + \text{ \_\_\_\_ HCl}$
- $\text{ \_\_\_\_ Cu} + 8 \text{ HNO}_3 = \text{ \_\_\_\_ Cu(NO}_3)_2 + 2 \text{ NO} + 4 \text{ H}_2\text{O}$
- $\text{ \_\_\_\_ BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 = 3 \text{ BaSO}_4 + \text{ \_\_\_\_ AlCl}_3$
- $\text{ \_\_\_\_ K} + \text{Br}_2 = \text{ \_\_\_\_ KBr}$
- $\text{B}_2\text{H}_6 + \text{ \_\_\_\_ H}_2\text{O} = 2 \text{ H}_3\text{BO}_3 + \text{ \_\_\_\_ H}_2$
- $\text{ \_\_\_\_ Ba(NO}_3\text{)}_2(\text{s}) = \text{ \_\_\_\_ BaO} + 4 \text{ NO}_2(\text{s}) + \text{O}_2(\text{s})$
- $2 \text{ Li} + \text{ \_\_\_\_ CO}_2 + 2 \text{ H}_2\text{O} = \text{ \_\_\_\_ LiHCO}_3 + \text{H}_2$
- $\text{ \_\_\_\_ KBr(aq)} + \text{Cl}_2(\text{g}) = \text{ \_\_\_\_ KCl(aq)} + \text{Br}_2(\text{l})$
- $\text{ \_\_\_\_ HNO}_3 + \text{CA(OH)}_2 = \text{CA(NO}_3)_2 + \text{ \_\_\_\_ H}_2\text{O}$
- $\text{ \_\_\_\_ C}_3\text{H}_3\text{O}_2 + 11 \text{ O}_2 = \text{ \_\_\_\_ CO}_2 + 6 \text{ H}_2\text{O}$
- $\text{ \_\_\_\_ CuSO}_4 + \text{CaCl}_2 = \text{ \_\_\_\_ CuCl} + \text{Ca(SO}_4)_2$



# ANSWERS

1.  $\text{Pb(s)} + 2 \text{AgNO}_3(\text{aq}) = \text{Pb(NO}_3)_2 + 2 \text{Ag(s)}$
2.  $\text{Fe}_2\text{O}_3 + 3 \text{CO} = 3 \text{CO}_2 + 2 \text{Fe}$
3.  $2 \text{Al(s)} + 2 \text{NaOH(aq)} + 6 \text{H}_2\text{O(l)} = 2 \text{Na(Al(OH)}_4\text{)(aq)} + 3 \text{H}_2(\text{g})$
4.  $2 \text{KMnO}_4 + 16 \text{HCl} = 2 \text{KCl} + 2 \text{MnCl}_2 + 8 \text{H}_2\text{O} + 5 \text{Cl}_2$
5.  $\text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{O}_2 = 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
6.  $3 \text{Zn} + 8 \text{HNO}_3 = 3 \text{Zn(NO}_3)_2 + 2 \text{NO} + 4 \text{H}_2\text{O}$
7.  $5 \text{Fe}^{(2+)} + \text{MNO}_4^{(-)} + 8 \text{H}^{(+)} = 5 \text{Fe}^{(3+)} + \text{MN}^{(2+)} + 4 \text{H}_2\text{O}$
8.  $2 \text{HCl} + \text{K}_2\text{SO}_3 = 2 \text{KCl} + \text{H}_2\text{O} + \text{SO}_2$
9.  $\text{TiCl}_4 + 2 \text{H}_2\text{O} = \text{TiO}_2 + 4 \text{HCl}$
10.  $2 \text{HSbCl}_4 + 3 \text{H}_2\text{S} = \text{Sb}_2\text{S}_3 + 8 \text{HCl}$
11.  $3 \text{Cu} + 8 \text{HNO}_3 = 3 \text{Cu(NO}_3)_2 + 2 \text{NO} + 4 \text{H}_2\text{O}$
12.  $3 \text{BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 = 3 \text{BaSO}_4 + 2 \text{AlCl}_3$
13.  $2 \text{K} + \text{Br}_2 = 2 \text{KBr}$
14.  $\text{B}_2\text{H}_6 + 6 \text{H}_2\text{O} = 2 \text{H}_3\text{BO}_3 + 6 \text{H}_2$
15.  $2 \text{Ba(NO}_3\text{(s))}_2\text{(s)} = 2 \text{BaO} + 4 \text{NO}_2\text{(s)} + \text{O}_2\text{(s)}$
16.  $2 \text{Li} + 2 \text{CO}_2 + 2 \text{H}_2\text{O} = 2 \text{LiHCO}_3 + \text{H}_2$
17.  $2 \text{KBr(aq)} + \text{Cl}_2(\text{g}) = 2 \text{KCl(aq)} + \text{Br}_2(\text{l})$
18.  $2 \text{HNO}_3 + \text{CA(OH)}_2 = \text{CA(NO}_3)_2 + 2 \text{H}_2\text{O}$
19.  $4 \text{C}_3\text{H}_8\text{O}_2 + 11 \text{O}_2 = 12 \text{CO}_2 + 6 \text{H}_2\text{O}$
20.  $2 \text{CuSO}_4 + \text{CaCl}_2 = 2 \text{CuCl} + \text{Ca(SO}_4)_2$



Thanks for downloading our free printable.

We have thousands of such resources in our blog for teachers and parents.

**[You can download them for free here!](#)**

### **Free Printables from Go Science Girls – Fair Usage Policy**

#### **You can ...**

- Download and save this free printable from [gosciencegirls.com](http://gosciencegirls.com) to your computer.
- Print this file and use it as many times as you want in your home, classrooms or for your library.
- Feel free to link our blog post where your visitors can find and download this printable for free.
- When you post online about this resource – please give due credit to [gosciencegirls.com](http://gosciencegirls.com)

#### **You Cannot ...**

- Access this file or download it from other sites apart from [gosciencegirls.com](http://gosciencegirls.com)
- Other websites cannot link to this pdf directly. If required, they are welcomed to link to the blog post from where this pdf can be downloaded.
- The ownership of this pdf rests with GoScienceGirls. No one can claim ownership for this file.
- You are not allowed to sell printed copies of this file to others.
- You are not allowed to store this file electronically and redistribute it (only personal use is allowed).

### **Further Questions?**

Feel free to email us at [contactgosciencegirls@gmail.com](mailto:contactgosciencegirls@gmail.com) for any further questions and suggestions. We would love to hear from you. We promise to respond back as soon as we can.