



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 7

- $2 \text{C}_{12}\text{H}_{26} + \text{O}_2 = 24 \text{CO}_2 + \text{H}_2\text{O}$
- $\text{C}_4\text{H}_8 + \text{O}_2 = \text{CO}_2 + 4 \text{H}_2\text{O}$
- $\text{CH}_4 + \text{O}_2 = \text{CO}_2 + \text{H}_2\text{O}$
- $\text{AgNO}_3 + \text{K}_4[\text{Fe}(\text{CN})_6] = \text{Ag}_4[\text{Fe}(\text{CN})_6] + \text{KNO}_3$
- $\text{CaCl}_2 + \text{Na}_2\text{SO}_4 = \text{Ca}(\text{SO}_4) + \text{NaCl}$
- $\text{Na}_3\text{PO}_4 + 3 \text{Ca}(\text{OH})_2 = \text{Ca}_3(\text{PO}_4)_2 + \text{NaOH}$
- $2 \text{Fe} + \text{CuO} = \text{Fe}_2\text{O}_3 + 3 \text{Cu}$
- $3 \text{Na}_2\text{CO}_3 + \text{Fe}(\text{NO}_3)_3 = 6 \text{NaNO}_3 + \text{Fe}_2(\text{CO}_3)_3$
- $\text{Pb}(\text{NO}_3)_2 = \text{Pb} + \text{NO}_2 + \text{O}_2$
- $\text{NH}_3 + 5 \text{O}_2 = \text{NO} + 6 \text{H}_2\text{O}$
- $\text{Cu} + \text{NH}_3 + 2 \text{H}_2\text{O} = \text{Cu}(\text{OH})_2 + 2 \text{NH}_4$
- $5 \text{H}_2\text{SO}_4 + \text{OH} + 2 \text{Cu}(\text{NH}_3)_4 = 2 \text{H}_2\text{O} + \text{Cu}_2\text{SO}_4 + (\text{NH}_4)_2\text{SO}_4$
- $4 \text{NH}_3 + \text{O}_2 = 4 \text{NO} + \text{H}_2\text{O}$
- $\text{CH}_3\text{CH}_2\text{CH}_3(\text{g}) + \text{O}_2(\text{g}) = \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O}(\text{g})$
- $2 \text{C}_4\text{H}_{10} + \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
- $4 \text{HCl} + \text{MnO}_2 = \text{MnCl}_2 + \text{H}_2\text{O} + \text{Cl}_2$
- $(\text{NH}_4)_2\text{S} + \text{Pb}(\text{NO}_2)_2 = \text{NH}_4\text{NO}_2 + \text{PbS}$
- $2 \text{HgO} = \text{Hg} + \text{O}_2$
- $\text{N}_2 + 3 \text{H}_2 = \text{NH}_3$
- $\text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + 2 \text{O}_2 + 4 \text{H}_2\text{O}$



ANSWERS

1. $2 \text{C}_{12}\text{H}_{26} + 37 \text{O}_2 = 24 \text{CO}_2 + 26 \text{H}_2\text{O}$
2. $\text{C}_4\text{H}_8 + 6 \text{O}_2 = 4 \text{CO}_2 + 4 \text{H}_2\text{O}$
3. $\text{CH}_4 + 2 \text{O}_2 = \text{CO}_2 + 2 \text{H}_2\text{O}$
4. $4 \text{AgNO}_3 + \text{K}_4[\text{Fe}(\text{CN})_6] = \text{Ag}_4[\text{Fe}(\text{CN})_6] + 4 \text{KNO}_3$
5. $\text{CaCl}_2 + \text{Na}_2\text{SO}_4 = \text{Ca}(\text{SO}_4) + 2 \text{NaCl}$
6. $2 \text{Na}_3\text{PO}_4 + 3 \text{Ca}(\text{OH})_2 = \text{Ca}_3(\text{PO}_4)_2 + 6 \text{NaOH}$
7. $2 \text{Fe} + 3 \text{CuO} = \text{Fe}_2\text{O}_3 + 3 \text{Cu}$
8. $3 \text{Na}_2\text{CO}_3 + 2 \text{Fe}(\text{NO}_3)_3 = 6 \text{NaNO}_3 + \text{Fe}_2(\text{CO}_3)_3$
9. $\text{Pb}(\text{NO}_3)_2 = \text{Pb} + 2 \text{NO}_2 + \text{O}_2$
10. $4 \text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + 6 \text{H}_2\text{O}$
11. $\text{Cu} + 2 \text{NH}_3 + 2 \text{H}_2\text{O} = \text{Cu}(\text{OH})_2 + 2 \text{NH}_4$
12. $5 \text{H}_2\text{SO}_4 + 2 \text{OH} + 2 \text{Cu}(\text{NH}_3)_4 = 2 \text{H}_2\text{O} + \text{Cu}_2\text{SO}_4 + 4 (\text{NH}_4)_2\text{SO}_4$
13. $4 \text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + 6 \text{H}_2\text{O}$
14. $\text{CH}_3\text{CH}_2\text{CH}_3(\text{g}) + 5 \text{O}_2(\text{g}) = 3 \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O}(\text{g})$
15. $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
16. $4 \text{HCl} + \text{MnO}_2 = \text{MnCl}_2 + 2 \text{H}_2\text{O} + \text{Cl}_2$
17. $(\text{NH}_4)_2\text{S} + \text{Pb}(\text{NO}_2)_2 = 2 \text{NH}_4\text{NO}_2 + \text{PbS}$
18. $2 \text{HgO} = 2 \text{Hg} + \text{O}_2$
19. $\text{N}_2 + 3 \text{H}_2 = 2 \text{NH}_3$
20. $2 \text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + 2 \text{O}_2 + 4 \text{H}_2\text{O}$



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