



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 67

- $2 \text{C}_4\text{H}_8\text{O}(\text{g}) + \text{ ______ } \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + \text{ ______ } \text{H}_2\text{O}(\text{g})$
- $\text{ ______ } \text{C}_6\text{H}_{14} + 19 \text{O}_2 = 12 \text{CO}_2 + \text{ ______ } \text{H}_2\text{O}$
- $\text{ ______ } \text{LiAlH}_4 + \text{AlCl}_3 = 4 \text{AlH}_3 + \text{ ______ } \text{LiCl}$
- $\text{ ______ } \text{FeO} + \text{O}_2 = \text{ ______ } \text{Fe}_2\text{O}_3$
- $\text{ ______ } \text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + \text{ ______ } \text{O}_2 + 4 \text{H}_2\text{O}$
- $4 \text{P} + \text{ ______ } \text{O}_2 = \text{ ______ } \text{P}_2\text{O}_5$
- $2 \text{RH} + \text{ ______ } \text{F}_2 = \text{ ______ } \text{RHF}_3$
- $\text{ ______ } \text{Cr} + 3 \text{H}_2\text{SO}_4 = \text{Cr}_2(\text{SO}_4)_3 + \text{ ______ } \text{H}_2$
- $\text{O}_2 + \text{ ______ } \text{CH}_4 = \text{ ______ } \text{CO} + 4 \text{H}_2$
- $\text{ ______ } \text{K} + 2 \text{H}_2\text{O} = \text{ ______ } \text{KOH} + \text{H}_2$
- $\text{V}_2\text{O}_5 + \text{ ______ } \text{H}_2 = \text{V}_2\text{O}_3 + \text{ ______ } \text{H}_2\text{O}$
- $3 \text{Na} + \text{ ______ } \text{HNO}_3 = \text{NO} + \text{ ______ } \text{H}_2\text{O} + 3 \text{NaNO}_3$
- $\text{ ______ } \text{Fe} + 4 \text{H}_2\text{O} = \text{Fe}_3\text{O}_4 + \text{ ______ } \text{H}_2$
- $4 \text{C}_2\text{H}_5\text{NH}_2(\text{g}) + \text{ ______ } \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + \text{ ______ } \text{H}_2\text{O}(\text{g}) + 2 \text{N}_2(\text{g})$
- $2 \text{C}_4\text{H}_{10} + \text{ ______ } \text{O}_2 = 8 \text{CO}_2 + \text{ ______ } \text{H}_2\text{O}$
- $(\text{NH}_4)_3\text{PO}_4 + \text{ ______ } \text{NaOH} = \text{Na}_3\text{PO}_4 + \text{ ______ } \text{NH}_3 + 3 \text{H}_2\text{O}$
- $\text{ ______ } \text{Ag} + 4 \text{HNO}_3 = \text{NO} + \text{ ______ } \text{H}_2\text{O} + 3 \text{AgNO}_3$
- $3 \text{P} + \text{ ______ } \text{HNO}_3 + 2 \text{H}_2\text{O} = 5 \text{NO} + \text{ ______ } \text{H}_3\text{PO}_4$
- $3 \text{Mg} + \text{ ______ } \text{FeCl}_3 = 3 \text{MgCl}_2 + \text{ ______ } \text{Fe}$
- $\text{C}_7\text{H}_{16}(\text{g}) + \text{ ______ } \text{O}_2(\text{g}) = 7 \text{CO}_2(\text{g}) + \text{ ______ } \text{H}_2\text{O}(\text{g})$



ANSWERS

1. $2 \text{C}_4\text{H}_8\text{O}(\text{g}) + 11 \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + 8 \text{H}_2\text{O}(\text{g})$
2. $2 \text{C}_6\text{H}_{14} + 19 \text{O}_2 = 12 \text{CO}_2 + 14 \text{H}_2\text{O}$
3. $3 \text{LiAlH}_4 + \text{AlCl}_3 = 4 \text{AlH}_3 + 3 \text{LiCl}$
4. $4 \text{FeO} + \text{O}_2 = 2 \text{Fe}_2\text{O}_3$
5. $2 \text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + 2 \text{O}_2 + 4 \text{H}_2\text{O}$
6. $4 \text{P} + 5 \text{O}_2 = 2 \text{P}_2\text{O}_5$
7. $2 \text{RH} + 3 \text{F}_2 = 2 \text{RHF}_3$
8. $2 \text{Cr} + 3 \text{H}_2\text{SO}_4 = \text{Cr}_2(\text{SO}_4)_3 + 3 \text{H}_2$
9. $\text{O}_2 + 2 \text{CH}_4 = 2 \text{CO} + 4 \text{H}_2$
10. $2 \text{K} + 2 \text{H}_2\text{O} = 2 \text{KOH} + \text{H}_2$
11. $\text{V}_2\text{O}_5 + 2 \text{H}_2 = \text{V}_2\text{O}_3 + 2 \text{H}_2\text{O}$
12. $3 \text{Na} + 4 \text{HNO}_3 = \text{NO} + 2 \text{H}_2\text{O} + 3 \text{NaNO}_3$
13. $3 \text{Fe} + 4 \text{H}_2\text{O} = \text{Fe}_3\text{O}_4 + 4 \text{H}_2$
14. $4 \text{C}_2\text{H}_5\text{NH}_2(\text{g}) + 15 \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + 14 \text{H}_2\text{O}(\text{g}) + 2 \text{N}_2(\text{g})$
15. $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
16. $(\text{NH}_4)_3\text{PO}_4 + 3 \text{NaOH} = \text{Na}_3\text{PO}_4 + 3 \text{NH}_3 + 3 \text{H}_2\text{O}$
17. $3 \text{Ag} + 4 \text{HNO}_3 = \text{NO} + 2 \text{H}_2\text{O} + 3 \text{AgNO}_3$
18. $3 \text{P} + 5 \text{HNO}_3 + 2 \text{H}_2\text{O} = 5 \text{NO} + 3 \text{H}_3\text{PO}_4$
19. $3 \text{Mg} + 2 \text{FeCl}_3 = 3 \text{MgCl}_2 + 2 \text{Fe}$
20. $\text{C}_7\text{H}_{16}(\text{g}) + 11 \text{O}_2(\text{g}) = 7 \text{CO}_2(\text{g}) + 8 \text{H}_2\text{O}(\text{g})$



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