



## BALANCE THE GIVEN CHEMICAL EQUATIONS

### Worksheet - 64

- $\text{Al}_2\text{S}_3 + \text{H}_2\text{O} = 2 \text{Al}(\text{OH})_3 + \text{H}_2\text{S}$
- $\text{C}_6\text{H}_{12}\text{O}_2 + 25 \text{O}_2 = \text{CO}_2 + 3 \text{H}_2\text{O}$
- $\text{S} + \text{HNO}_3 = \text{H}_2\text{SO}_4 + \text{NO}_2 + 2 \text{H}_2\text{O}$
- $\text{CH}_3(\text{CH}_2)_6\text{CH}_3 + 25 \text{O}_2 = \text{CO}_2 + 18 \text{H}_2\text{O}$
- $\text{Cr}_2\text{O}_3 + \text{Cl}_2 + 3 \text{C} = \text{CrCl}_3 + 3 \text{CO}$
- $\text{C}_2\text{H}_6 + 7 \text{O}_2 = 6 \text{H}_2\text{O} + \text{CO}_2$
- $2 \text{CH}_3\text{OH} + \text{O}_2 = \text{CO}_2 + 4 \text{H}_2\text{O}$
- $(\text{NH}_4)_2\text{Cr}_2\text{O}_7 = \text{NH}_3 + 14 \text{H}_2\text{O} + 5 \text{Cr}_2\text{O}_3 + 6 \text{NO}$
- $\text{NaCl} + \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + \text{HCl}$
- $2 \text{C}_5\text{H}_{10}\text{O}_2(\text{l}) + \text{O}_2(\text{g}) = 10 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
- $\text{CH}_4 + \text{O}_2 = \text{CO}_2 + \text{H}_2\text{O}$
- $\text{Cr}_2\text{O}_3 + \text{Cl}_2 + 3 \text{C} = 2 \text{CrCl}_3 + \text{CO}$
- $\text{HNO}_3 + \text{H}_2\text{S} = \text{NO}_2 + \text{H}_2\text{SO}_4 + 4 \text{H}_2\text{O}$
- $\text{GaF}_3 + \text{Cs} = \text{CsF} + \text{Ga}$
- $\text{C}_9\text{H}_{18}\text{O} + \text{O}_2 = 9 \text{CO}_2 + \text{H}_2\text{O}$
- $3 \text{Cu} + \text{HNO}_3 = 3 \text{Cu}(\text{NO}_3)_2 + \text{NO} + 4 \text{H}_2\text{O}$
- $\text{Na} + 2 \text{H}_2\text{O} = \text{NaOH} + \text{H}_2$
- $2 \text{Al} + \text{HCl} = \text{AlCl}_3 + 3 \text{H}_2$
- $4 \text{NH}_3 + \text{O}_2 = 4 \text{NO} + \text{H}_2\text{O}$
- $\text{FeCl}_2 + \text{H}_2\text{O}_2 + 2 \text{HCl} = \text{FeCl}_3 + 2 \text{H}_2\text{O}$



# ANSWERS

1.  $\text{Al}_2\text{S}_3 + 6 \text{H}_2\text{O} = 2 \text{Al}(\text{OH})_3 + 3 \text{H}_2\text{S}$
2.  $5 \text{C}_6\text{H}_{12}\text{O}_2 + 25 \text{O}_2 = 30 \text{CO}_2 + 3 \text{H}_2\text{O}$
3.  $\text{S} + 6 \text{HNO}_3 = \text{H}_2\text{SO}_4 + 6 \text{NO}_2 + 2 \text{H}_2\text{O}$
4.  $2 \text{CH}_3(\text{CH}_2)_6\text{CH}_3 + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
5.  $\text{Cr}_2\text{O}_3 + 3 \text{Cl}_2 + 3 \text{C} = 2 \text{CrCl}_3 + 3 \text{CO}$
6.  $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 = 6 \text{H}_2\text{O} + 4 \text{CO}_2$
7.  $2 \text{CH}_3\text{OH} + 3 \text{O}_2 = 2 \text{CO}_2 + 4 \text{H}_2\text{O}$
8.  $5 (\text{NH}_4)_2\text{Cr}_2\text{O}_7 = 4 \text{NH}_3 + 14 \text{H}_2\text{O} + 5 \text{Cr}_2\text{O}_3 + 6 \text{NO}$
9.  $2 \text{NaCl} + \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + 2 \text{HCl}$
10.  $2 \text{C}_5\text{H}_{10}\text{O}_2(\text{l}) + 13 \text{O}_2(\text{g}) = 10 \text{CO}_2(\text{g}) + 10 \text{H}_2\text{O}(\text{g})$
11.  $\text{CH}_4 + 2 \text{O}_2 = \text{CO}_2 + 2 \text{H}_2\text{O}$
12.  $\text{Cr}_2\text{O}_3 + 3 \text{Cl}_2 + 3 \text{C} = 2 \text{CrCl}_3 + 3 \text{CO}$
13.  $8 \text{HNO}_3 + \text{H}_2\text{S} = 8 \text{NO}_2 + \text{H}_2\text{SO}_4 + 4 \text{H}_2\text{O}$
14.  $\text{GaF}_3 + 3 \text{Cs} = 3 \text{CsF} + \text{Ga}$
15.  $\text{C}_9\text{H}_{18}\text{O} + 13 \text{O}_2 = 9 \text{CO}_2 + 9 \text{H}_2\text{O}$
16.  $3 \text{Cu} + 8 \text{HNO}_3 = 3 \text{Cu}(\text{NO}_3)_2 + 2 \text{NO} + 4 \text{H}_2\text{O}$
17.  $2 \text{Na} + 2 \text{H}_2\text{O} = 2 \text{NaOH} + \text{H}_2$
18.  $2 \text{Al} + 6 \text{HCl} = 2 \text{AlCl}_3 + 3 \text{H}_2$
19.  $4 \text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + 6 \text{H}_2\text{O}$
20.  $2 \text{FeCl}_2 + \text{H}_2\text{O}_2 + 2 \text{HCl} = 2 \text{FeCl}_3 + 2 \text{H}_2\text{O}$



Thanks for downloading our free printable.

We have thousands of such resources in our blog for teachers and parents.

**[You can download them for free here!](#)**

### **Free Printables from Go Science Girls – Fair Usage Policy**

#### **You can ...**

- Download and save this free printable from [gosciencegirls.com](http://gosciencegirls.com) to your computer.
- Print this file and use it as many times as you want in your home, classrooms or for your library.
- Feel free to link our blog post where your visitors can find and download this printable for free.
- When you post online about this resource – please give due credit to [gosciencegirls.com](http://gosciencegirls.com)

#### **You Cannot ...**

- Access this file or download it from other sites apart from [gosciencegirls.com](http://gosciencegirls.com)
- Other websites cannot link to this pdf directly. If required, they are welcomed to link to the blog post from where this pdf can be downloaded.
- The ownership of this pdf rests with GoScienceGirls. No one can claim ownership for this file.
- You are not allowed to sell printed copies of this file to others.
- You are not allowed to store this file electronically and redistribute it (only personal use is allowed).

### **Further Questions?**

Feel free to email us at [contactgosciencegirls@gmail.com](mailto:contactgosciencegirls@gmail.com) for any further questions and suggestions. We would love to hear from you. We promise to respond back as soon as we can.