



## BALANCE THE GIVEN CHEMICAL EQUATIONS

### Worksheet - 19

1.  $2 \text{K} + \text{HCl} = \text{H}_2 + \text{KCl}$
2.  $\text{P} + 5 \text{O}_2 = \text{P}_2\text{O}_5$
3.  $\text{Li} + 10 \text{HNO}_3 = \text{LiNO}_3 + \text{NH}_4\text{NO}_3 + 3 \text{H}_2\text{O}$
4.  $\text{Mg} + \text{O}_2 = \text{MgO}$
5.  $\text{P}_2\text{O}_5 + \text{NaOH} = 3 \text{H}_2\text{O} + \text{Na}_3\text{PO}_4$
6.  $\text{S}_8 + \text{O}_2 = \text{SO}_3$
7.  $\text{KClO}_3 = \text{KCl} + 3 \text{O}_2$
8.  $\text{Sb} + 3 \text{O}_2 = \text{Sb}_4\text{O}_6$
9.  $\text{CuS} + 4 \text{HNO}_3 = \text{Cu}(\text{NO}_3)_2 + 3 \text{S} + \text{NO} + \text{H}_2\text{O}$
10.  $\text{FeS}_2 + 10 \text{HNO}_3 = \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2\text{SO}_4 + \text{NO} + 4 \text{H}_2\text{O}$
11.  $\text{HNO}_3 + \text{Ba}(\text{OH})_2 = \text{Ba}(\text{NO}_3)_2 + \text{H}_2\text{O}$
12.  $\text{Br}_2 + 3 \text{O}_2 = \text{Br}_2\text{O}_3$
13.  $2 \text{N}_2 + \text{O}_2 = \text{N}_2\text{O}_5$
14.  $2 \text{Fe}_2\text{O}_3 + \text{C} = 4 \text{Fe} + \text{CO}_2$
15.  $\text{SO}_3 = \text{SO}_2 + \text{O}_2$
16.  $2 \text{CH}_3(\text{CH}_2)_6\text{CH}_3 + \text{O}_2 = 16 \text{CO}_2 + \text{H}_2\text{O}$
17.  $\text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + \text{H}_2\text{O}$
18.  $\text{C}_3\text{H}_8 + \text{O}_2 = 3 \text{CO}_2 + \text{H}_2\text{O}$
19.  $2 \text{K}_2\text{S} + \text{O}_2 = \text{K}_2\text{SO}_3$
20.  $\text{CH}_3\text{OH}(\text{l}) + 3 \text{O}_2(\text{g}) = \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O}(\text{g})$



# ANSWERS

1.  $2 \text{K} + 2 \text{HCl} = \text{H}_2 + 2 \text{KCl}$
2.  $4 \text{P} + 5 \text{O}_2 = 2 \text{P}_2\text{O}_5$
3.  $8 \text{Li} + 10 \text{HNO}_3 = 8 \text{LiNO}_3 + \text{NH}_4\text{NO}_3 + 3 \text{H}_2\text{O}$
4.  $2 \text{Mg} + \text{O}_2 = 2 \text{MgO}$
5.  $\text{P}_2\text{O}_5 + 6 \text{NaOH} = 3 \text{H}_2\text{O} + 2 \text{Na}_3\text{PO}_4$
6.  $\text{S}_8 + 12 \text{O}_2 = 8 \text{SO}_3$
7.  $2 \text{KClO}_3 = 2 \text{KCl} + 3 \text{O}_2$
8.  $4 \text{Sb} + 3 \text{O}_2 = \text{Sb}_4\text{O}_6$
9.  $3 \text{CuS} + 4 \text{HNO}_3 = 3 \text{Cu}(\text{NO}_3) + 3 \text{S} + \text{NO} + 2 \text{H}_2\text{O}$
10.  $2 \text{FeS}_2 + 10 \text{HNO}_3 = \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2\text{SO}_4 + 10 \text{NO} + 4 \text{H}_2\text{O}$
11.  $2 \text{HNO}_3 + \text{Ba}(\text{OH})_2 = \text{Ba}(\text{NO}_3)_2 + 2 \text{H}_2\text{O}$
12.  $2 \text{Br}_2 + 3 \text{O}_2 = 2 \text{Br}_2\text{O}_3$
13.  $2 \text{N}_2 + 5 \text{O}_2 = 2 \text{N}_2\text{O}_5$
14.  $2 \text{Fe}_2\text{O}_3 + 3 \text{C} = 4 \text{Fe} + 3 \text{CO}_2$
15.  $2 \text{SO}_3 = 2 \text{SO}_2 + \text{O}_2$
16.  $2 \text{CH}_3(\text{CH}_2)_6\text{CH}_3 + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
17.  $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
18.  $\text{C}_3\text{H}_8 + 5 \text{O}_2 = 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
19.  $2 \text{K}_2\text{S} + 3 \text{O}_2 = 2 \text{K}_2\text{SO}_3$
20.  $2 \text{CH}_3\text{OH}(\text{l}) + 3 \text{O}_2(\text{g}) = 2 \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O}(\text{g})$



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