



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 53

- $2 \text{C}_6\text{H}_{14}\text{O}_4 + \text{O}_2 = \text{CO}_2 + 14 \text{H}_2\text{O}$
- $\text{KMnO}_4 + 5 \text{H}_2\text{C}_2\text{O}_4 + 3 \text{H}_2\text{SO}_4 = \text{MnSO}_4 + \text{K}_2\text{SO}_4 + 10 \text{CO}_2 + 8 \text{H}_2\text{O}$
- $\text{C}_2\text{H}_6 + 7 \text{O}_2 = 4 \text{CO}_2 + \text{H}_2\text{O}$
- $3 \text{C} + \text{Fe}_2\text{O}_3 = \text{Fe} + 3 \text{CO}_2$
- $2 \text{KMnO}_4 + \text{HCl} = 2 \text{MnCl}_2 + \text{Cl}_2 + 8 \text{H}_2\text{O} + 2 \text{KCl}$
- $2 \text{C}_2\text{H}_5\text{OH} + \text{Na} = \text{C}_2\text{H}_5\text{ONa} + \text{H}_2$
- $\text{Al}(\text{MnO}_4)_3 + 3 \text{Hg}(\text{SO}_4) = \text{Al}_2(\text{SO}_4)_3 + \text{Hg}(\text{MnO}_4)_2$
- $25 \text{O}_2 + \text{C}_8\text{H}_{18} = \text{CO}_2 + 18 \text{H}_2\text{O}$
- $\text{Al}(\text{NO}_3)_3 + 3 \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + \text{HNO}_3$
- $2 \text{C}_8\text{H}_{18} + \text{O}_2 = \text{CO}_2 + 18 \text{H}_2\text{O}$
- $\text{K}(\text{s}) + \text{N}_2(\text{g}) = \text{K}_3\text{N}(\text{s})$
- $2 \text{Na} + \text{H}_2\text{O} = \text{NaOH} + \text{H}_2$
- $\text{C}_8\text{H}_{18} + 25 \text{O}_2 = \text{CO}_2 + 18 \text{H}_2\text{O}$
- $\text{C}_4\text{H}_8(\text{g}) + \text{O}_2(\text{g}) = 4 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
- $\text{CH}_3(\text{CH}_2)_3\text{CH}_3 + \text{O}_2 = 5 \text{CO}_2 + \text{H}_2\text{O}$
- $\text{C}_2\text{H}_6 + 7 \text{O}_2 = \text{CO}_2 + 6 \text{H}_2\text{O}$
- $2 \text{C}_8\text{H}_{18} + \text{O}_2 = 16 \text{CO}_2 + \text{H}_2\text{O}$
- $\text{C}_8\text{H}_{18} + 25 \text{O}_2 = \text{CO}_2 + 18 \text{H}_2\text{O}$
- $2 \text{C}_3\text{H}_6(\text{g}) + \text{O}_2(\text{g}) = 6 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
- $\text{HBr} + \text{Ba}(\text{OH})_2 = \text{H}_2\text{O} + \text{BaBr}_2$



ANSWERS

1. $2 \text{C}_6\text{H}_{14}\text{O}_4 + 15 \text{O}_2 = 12 \text{CO}_2 + 14 \text{H}_2\text{O}$
2. $2 \text{KMNO}_4 + 5 \text{H}_2\text{C}_2\text{O}_4 + 3 \text{H}_2\text{SO}_4 = 2 \text{MNSO}_4 + \text{K}_2\text{SO}_4 + 10 \text{CO}_2 + 8 \text{H}_2\text{O}$
3. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
4. $3 \text{C} + 2 \text{Fe}_2\text{O}_3 = 4 \text{Fe} + 3 \text{CO}_2$
5. $2 \text{KMnO}_4 + 16 \text{HCl} = 2 \text{MnCl}_2 + 5 \text{Cl}_2 + 8 \text{H}_2\text{O} + 2 \text{KCl}$
6. $2 \text{C}_2\text{H}_5\text{OH} + 2 \text{Na} = 2 \text{C}_2\text{H}_5\text{ONa} + \text{H}_2$
7. $2 \text{Al}(\text{MnO}_4)_3 + 3 \text{Hg}(\text{SO}_4) = \text{Al}_2(\text{SO}_4)_3 + 3 \text{Hg}(\text{MnO}_4)_2$
8. $25 \text{O}_2 + 2 \text{C}_8\text{H}_{18} = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
9. $2 \text{Al}(\text{NO}_3)_3 + 3 \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 6 \text{HNO}_3$
10. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
11. $6 \text{K}(\text{s}) + \text{N}_2(\text{g}) = 2 \text{K}_3\text{N}(\text{s})$
12. $2 \text{Na} + 2 \text{H}_2\text{O} = 2 \text{NaOH} + \text{H}_2$
13. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
14. $\text{C}_4\text{H}_8(\text{g}) + 6 \text{O}_2(\text{g}) = 4 \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O}(\text{g})$
15. $\text{CH}_3(\text{CH}_2)_3\text{CH}_3 + 8 \text{O}_2 = 5 \text{CO}_2 + 6 \text{H}_2\text{O}$
16. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
17. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
18. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
19. $2 \text{C}_3\text{H}_6(\text{g}) + 9 \text{O}_2(\text{g}) = 6 \text{CO}_2(\text{g}) + 6 \text{H}_2\text{O}(\text{g})$
20. $2 \text{HBr} + \text{Ba}(\text{OH})_2 = 2 \text{H}_2\text{O} + \text{BaBr}_2$



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