



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 48

1. $\text{Si}_4\text{H}_{10} + 13 \text{O}_2 = \text{SiO}_2 + 10 \text{H}_2\text{O}$
2. $\text{HClO}_3 + \text{HCl} = 3 \text{Cl}_2 + \text{H}_2\text{O}$
3. $20 \text{MnO}_4^{-} + \text{NO}_2^{-} + 60 \text{H}^{+} = \text{Mn}^{2+} + 80 \text{NO}_3^{-} + 3 \text{H}_2\text{O}$
4. $\text{KClO}_3 = \text{KCl} + 3 \text{O}_2$
5. $4 \text{NH}_3 + \text{O}_2 = 4 \text{NO}_2 + \text{H}_2\text{O}$
6. $2 \text{Al} + \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + \text{H}_2$
7. $2 \text{Bi}(\text{OH})_3 + \text{Na}_2\text{SnO}_2 = \text{Na}_2\text{SnO}_3 + 2 \text{Bi} + 3 \text{H}_2\text{O}$
8. $\text{H}_2\text{SO}_4 + \text{NaOH} = \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
9. $2 \text{C}_8\text{H}_{18} + \text{O}_2 = \text{CO}_2 + 18 \text{H}_2\text{O}$
10. $\text{C} + 2 \text{SO}_2 = \text{CS}_2 + \text{CO}$
11. $\text{Al}_2\text{O}_3 + \text{K}_2\text{S}_2\text{O}_7 = \text{Al}_2(\text{SO}_4)_3 + \text{K}_2\text{SO}_4$
12. $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CrO}_4 + \text{H}_3\text{O} = \text{H}_2\text{CrO}_3 + \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
13. $\text{C}_4\text{H}_8 + \text{O}_2 = \text{CO}_2 + 4 \text{H}_2\text{O}$
14. $\text{C}_6\text{H}_{12} + \text{O}_2 = \text{CO}_2 + 6 \text{H}_2\text{O}$
15. $\text{C}_9\text{H}_{20} + \text{O}_2 = \text{CO}_2 + \text{H}_2\text{O}$
16. $2 \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) = \text{CO}_2(\text{g}) + 6 \text{H}_2\text{O}(\text{g})$
17. $\text{C}_6\text{H}_{12} + \text{O}_2 = \text{CO}_2 + 6 \text{H}_2\text{O}$
18. $\text{Ba}(\text{OH})_2 + \text{HClO}_3 = \text{Ba}(\text{ClO}_3)_2 + \text{H}_2\text{O}$
19. $\text{AgNO}_3 + \text{CaCl}_2 = \text{AgCl} + \text{Ca}(\text{NO}_3)_2$
20. $\text{C}_3\text{H}_8 + 7 \text{O}_2 = \text{CO} + 8 \text{H}_2\text{O}$



ANSWERS

1. $2 \text{Si}_4\text{H}_{10} + 13 \text{O}_2 = 8 \text{SiO}_2 + 10 \text{H}_2\text{O}$
2. $\text{HClO}_3 + 5 \text{HCl} = 3 \text{Cl}_2 + 3 \text{H}_2\text{O}$
3. $20 \text{MnO}_4^{\{-\}} + 80 \text{NO}_2^{\{-\}} + 60 \text{H}^{\{+\}} = 20 \text{Mn}^{\{2+\}} + 80 \text{NO}_3^{\{-\}} + 3 \text{H}_2\text{O}$
4. $2 \text{KClO}_3 = 2 \text{KCl} + 3 \text{O}_2$
5. $4 \text{NH}_3 + 7 \text{O}_2 = 4 \text{NO}_2 + 6 \text{H}_2\text{O}$
6. $2 \text{Al} + 3 \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 3 \text{H}_2$
7. $2 \text{Bi}(\text{OH})_3 + 3 \text{Na}_2\text{SnO}_2 = 3 \text{Na}_2\text{SnO}_3 + 2 \text{Bi} + 3 \text{H}_2\text{O}$
8. $\text{H}_2\text{SO}_4 + 2 \text{NaOH} = \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
9. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
10. $5 \text{C} + 2 \text{SO}_2 = \text{CS}_2 + 4 \text{CO}$
11. $\text{Al}_2\text{O}_3 + 3 \text{K}_2\text{S}_2\text{O}_7 = \text{Al}_2(\text{SO}_4)_3 + 3 \text{K}_2\text{SO}_4$
12. $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CrO}_4 + 2 \text{H}_3\text{O} = \text{H}_2\text{CrO}_3 + \text{Na}_2\text{SO}_4 + 3 \text{H}_2\text{O}$
13. $\text{C}_4\text{H}_8 + 6 \text{O}_2 = 4 \text{CO}_2 + 4 \text{H}_2\text{O}$
14. $\text{C}_6\text{H}_{12} + 9 \text{O}_2 = 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
15. $\text{C}_9\text{H}_{20} + 9 \text{O}_2 = 9 \text{CO}_2 + \text{H}_2\text{O}$
16. $2 \text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) = 4 \text{CO}_2(\text{g}) + 6 \text{H}_2\text{O}(\text{g})$
17. $\text{C}_6\text{H}_{12} + 9 \text{O}_2 = 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
18. $\text{Ba}(\text{OH})_2 + 2 \text{HClO}_3 = \text{Ba}(\text{ClO}_3)_2 + 2 \text{H}_2\text{O}$
19. $2 \text{AgNO}_3 + \text{CaCl}_2 = 2 \text{AgCl} + \text{Ca}(\text{NO}_3)_2$
20. $2 \text{C}_3\text{H}_8 + 7 \text{O}_2 = 6 \text{CO} + 8 \text{H}_2\text{O}$



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