



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 26

1. ____ $\text{AgNO}_3(\text{aq}) + (\text{NH}_4)_2\text{SO}_4(\text{aq}) =$ ____ $\text{NH}_4\text{NO}_3(\text{aq}) + \text{Ag}_2\text{SO}_4(\text{s})$
2. $\text{N}_2\text{O}_5 + \text{H}_2\text{O} =$ ____ HNO_3
3. ____ $\text{NaOH} + \text{C}_2\text{H}_2\text{O}_4 = \text{Na}_2\text{C}_2\text{O}_4 +$ ____ H_2O
4. ____ $\text{Na} + \text{Mg}_3(\text{PO}_4)_2 =$ ____ $\text{Na}_3\text{PO}_4 + 3 \text{Mg}$
5. $\text{Al}_2(\text{SO}_4)_3 +$ ____ $\text{H}_3\text{PO}_4 = 2 \text{AlPO}_4 +$ ____ H_2SO_4
6. ____ $\text{Zn} + 4 \text{H}_2\text{SO}_4 =$ ____ $\text{ZnSO}_4 + 4 \text{H}_2\text{O} + \text{S}$
7. $\text{H}_3\text{PO}_4 +$ ____ $\text{NaOH} = \text{Na}_3\text{PO}_4 +$ ____ H_2O
8. $2 \text{KMnO}_4 +$ ____ $\text{H}_2\text{SO}_4 = \text{K}_2\text{SO}_4 + 2 \text{MnSO}_4 +$ ____ $\text{H}_2\text{O} + 5 \text{O}$
9. $2 \text{FeCl}_3 + \text{Cu} =$ ____ $\text{FeCl}_2 + \text{CuCl}_2$
10. ____ $\text{K}(\text{ClO}_3) = 2 \text{KCl} + 3 \text{O}_2$
11. $3 \text{CH}_3(\text{CH}_2)_{16}\text{COONa}(\text{s}) + \text{FeCl}_3(\text{s}) = \text{Fe}[\text{CH}_3(\text{CH}_2)_{16}\text{COO}]_3(\text{aq}) +$ ____ $\text{NaCl}(\text{aq})$
12. $2 \text{Al} +$ ____ $\text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 +$ ____ H_2
13. $\text{K}_2\text{O}(\text{s}) + \text{H}_2\text{O}(\text{l}) =$ ____ $\text{KOH}(\text{aq})$
14. ____ $\text{KI} + \text{Pb}(\text{NO}_3)_2 =$ ____ $\text{KNO}_3 + \text{PbI}_2$
15. $\text{Mg}_3\text{N}_2(\text{s}) +$ ____ $\text{H}_2\text{SO}_4(\text{aq}) = 3 \text{MgSO}_4(\text{aq}) + (\text{NH}_4)_2\text{SO}_4(\text{aq})$
16. $2 \text{Al} +$ ____ $\text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 3 \text{H}_2$
17. $2 \text{PBS} +$ ____ $\text{O}_2 = 2 \text{PBO} +$ ____ SO_2
18. $2 \text{NO}(\text{g}) +$ ____ $\text{H}_2(\text{g}) = \text{N}_2(\text{g}) +$ ____ $\text{H}_2\text{O}(\text{l})$
19. $\text{Sn} +$ ____ $\text{NaOH} = \text{Na}_2\text{SnO}_2 + \text{H}_2$
20. $3 \text{Ba}(\text{OH})_2 +$ ____ $\text{H}_3\text{PO}_4 = \text{Ba}_3(\text{PO}_4)_2 +$ ____ H_2O



ANSWERS

1. $2 \text{AgNO}_3(\text{aq}) + (\text{NH}_4)_2\text{SO}_4(\text{aq}) = 2 \text{NH}_4\text{NO}_3(\text{aq}) + \text{Ag}_2\text{SO}_4(\text{s})$
2. $\text{N}_2\text{O}_5 + \text{H}_2\text{O} = 2 \text{HNO}_3$
3. $2 \text{NaOH} + \text{C}_2\text{H}_2\text{O}_4 = \text{Na}_2\text{C}_2\text{O}_4 + 2 \text{H}_2\text{O}$
4. $6 \text{Na} + \text{Mg}_3(\text{PO}_4)_2 = 2 \text{Na}_3\text{PO}_4 + 3 \text{Mg}$
5. $\text{Al}_2(\text{SO}_4)_3 + 2 \text{H}_3\text{PO}_4 = 2 \text{AlPO}_4 + 3 \text{H}_2\text{SO}_4$
6. $3 \text{Zn} + 4 \text{H}_2\text{SO}_4 = 3 \text{ZnSO}_4 + 4 \text{H}_2\text{O} + \text{S}$
7. $\text{H}_3\text{PO}_4 + 3 \text{NaOH} = \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
8. $2 \text{KMnO}_4 + 3 \text{H}_2\text{SO}_4 = \text{K}_2\text{SO}_4 + 2 \text{MnSO}_4 + 3 \text{H}_2\text{O} + 5 \text{O}$
9. $2 \text{FeCl}_3 + \text{Cu} = 2 \text{FeCl}_2 + \text{CuCl}_2$
10. $2 \text{K}(\text{ClO}_3) = 2 \text{KCl} + 3 \text{O}_2$
11. $3 \text{CH}_3(\text{CH}_2)_{16}\text{COONa}(\text{s}) + \text{FeCl}_3(\text{s}) = \text{Fe}[\text{CH}_3(\text{CH}_2)_{16}\text{COO}]_3(\text{aq}) + 3 \text{NaCl}(\text{aq})$
12. $2 \text{Al} + 3 \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 3 \text{H}_2$
13. $\text{K}_2\text{O}(\text{s}) + \text{H}_2\text{O}(\text{l}) = 2 \text{KOH}(\text{aq})$
14. $2 \text{KI} + \text{Pb}(\text{NO}_3)_2 = 2 \text{KNO}_3 + \text{PbI}_2$
15. $\text{Mg}_3\text{N}_2(\text{s}) + 4 \text{H}_2\text{SO}_4(\text{aq}) = 3 \text{MgSO}_4(\text{aq}) + (\text{NH}_4)_2\text{SO}_4(\text{aq})$
16. $2 \text{Al} + 3 \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 3 \text{H}_2$
17. $2 \text{PBS} + 3 \text{O}_2 = 2 \text{PBO} + 2 \text{SO}_2$
18. $2 \text{NO}(\text{g}) + 2 \text{H}_2(\text{g}) = \text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$
19. $\text{Sn} + 2 \text{NaOH} = \text{Na}_2\text{SnO}_2 + \text{H}_2$
20. $3 \text{Ba}(\text{OH})_2 + 2 \text{H}_3\text{PO}_4 = \text{Ba}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$



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