



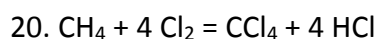
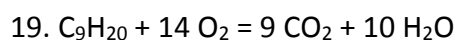
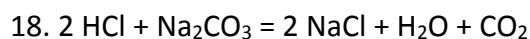
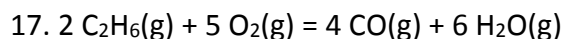
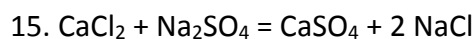
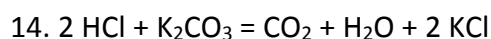
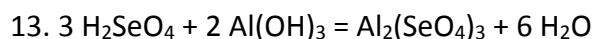
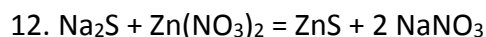
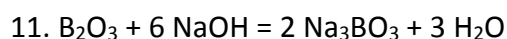
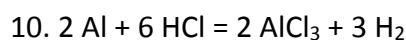
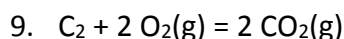
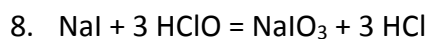
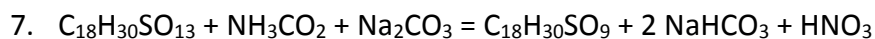
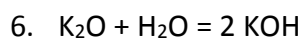
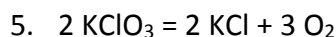
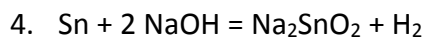
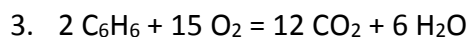
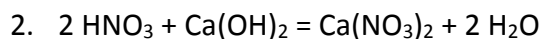
## BALANCE THE GIVEN CHEMICAL EQUATIONS

### Worksheet - 16

- $14 \text{ HCl} + \text{K}_2\text{Cr}_2\text{O}_7 + \text{ \_\_\_\_ } \text{Na}_2\text{S} = 2 \text{ CrCl}_3 + 7 \text{ H}_2\text{O} + 3 \text{ S} + \text{ \_\_\_\_ } \text{KCl} + 6 \text{ NaCl}$
- $\text{ \_\_\_\_ } \text{HNO}_3 + \text{Ca(OH)}_2 = \text{Ca(NO}_3)_2 + \text{ \_\_\_\_ } \text{H}_2\text{O}$
- $2 \text{ C}_6\text{H}_6 + \text{ \_\_\_\_ } \text{O}_2 = 12 \text{ CO}_2 + \text{ \_\_\_\_ } \text{H}_2\text{O}$
- $\text{Sn} + \text{ \_\_\_\_ } \text{NaOH} = \text{Na}_2\text{SnO}_2 + \text{H}_2$
- $\text{ \_\_\_\_ } \text{KClO}_3 = \text{ \_\_\_\_ } \text{KCl} + 3 \text{ O}_2$
- $\text{K}_2\text{O} + \text{H}_2\text{O} = \text{ \_\_\_\_ } \text{KOH}$
- $\text{C}_{18}\text{H}_{30}\text{SO}_{13} + \text{NH}_3\text{CO}_2 + \text{Na}_2\text{CO}_3 = \text{C}_{18}\text{H}_{30}\text{SO}_9 + \text{ \_\_\_\_ } \text{NaHCO}_3 + \text{HNO}_3$
- $\text{NaI} + \text{ \_\_\_\_ } \text{HClO} = \text{NaIO}_3 + 3 \text{ HCl}$
- $\text{C}_2 + \text{ \_\_\_\_ } \text{O}_2(\text{g}) = \text{ \_\_\_\_ } \text{CO}_2(\text{g})$
- $2 \text{ Al} + \text{ \_\_\_\_ } \text{HCl} = 2 \text{ AlCl}_3 + \text{ \_\_\_\_ } \text{H}_2$
- $\text{B}_2\text{O}_3 + \text{ \_\_\_\_ } \text{NaOH} = \text{ \_\_\_\_ } \text{Na}_3\text{BO}_3 + 3 \text{ H}_2\text{O}$
- $\text{Na}_2\text{S} + \text{Zn(NO}_3)_2 = \text{ZnS} + \text{ \_\_\_\_ } \text{NaNO}_3$
- $3 \text{ H}_2\text{SeO}_4 + 2 \text{ Al(OH)}_3 = \text{Al}_2(\text{SeO}_4)_3 + \text{ \_\_\_\_ } \text{H}_2\text{O}$
- $\text{ \_\_\_\_ } \text{HCl} + \text{K}_2\text{CO}_3 = \text{CO}_2 + \text{H}_2\text{O} + \text{ \_\_\_\_ } \text{KCl}$
- $\text{CaCl}_2 + \text{Na}_2\text{SO}_4 = \text{CaSO}_4 + \text{ \_\_\_\_ } \text{NaCl}$
- $\text{H}_3\text{PO}_4 + \text{ \_\_\_\_ } \text{NH}_3 = (\text{NH}_4)_3\text{PO}_4$
- $2 \text{ C}_2\text{H}_6(\text{g}) + \text{ \_\_\_\_ } \text{O}_2(\text{g}) = 4 \text{ CO}(\text{g}) + 6 \text{ H}_2\text{O}(\text{g})$
- $\text{ \_\_\_\_ } \text{HCl} + \text{Na}_2\text{CO}_3 = \text{ \_\_\_\_ } \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- $\text{C}_9\text{H}_{20} + \text{ \_\_\_\_ } \text{O}_2 = \text{ \_\_\_\_ } \text{CO}_2 + 10 \text{ H}_2\text{O}$
- $\text{CH}_4 + \text{ \_\_\_\_ } \text{Cl}_2 = \text{CCl}_4 + \text{ \_\_\_\_ } \text{HCl}$



# ANSWERS





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